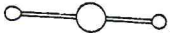
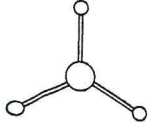
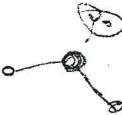
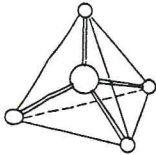
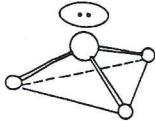
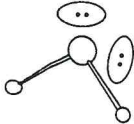
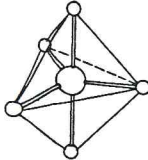
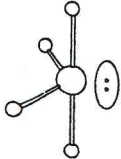
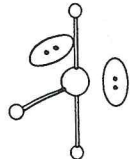
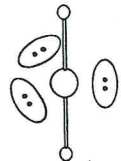
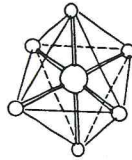
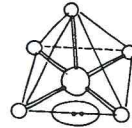
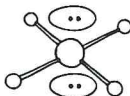


Número estérico	Números de pares isolados	Geometria molecular	
2	0	Linear	
3	0	* Plana trigonal	
3	1	FORMA V	
4	0	Tetraédrico	
4	1	Pirâmide trigonal	
4	2	Angular	
5	0	Bipirâmide trigonal	
5	1	Gangorra	

(continua na próxima pág.)

TABELA 9-2 (continuação)

Número estérico	Números de pares isolados	Geometria molecular
5	2	Forma T 
5	3	Linear 
6	0	Octaédrica 
6	1	Pirâmide tetragonal 
6	2	Quadrado planar 

COMENTÁRIO ADICIONAL

Um erro comum é confundir a geometria de um conjunto completo de pares de elétrons (compartilhados mais os isolados) com a forma da molécula. No AsCl_3 , por exemplo, a orientação dos pares de elétrons ao redor do átomo de As é tetraédrica (veja Fig. 9-3). Mas dizer que a